Test Solutions and Solubility SCH3U

Name:

A Assessment of Learning (Product) 1.5 14 9. 10

Multiple Choice. Circle the most correct answer. (15K, 5T)

How many of the following salts can be expected to be insoluble in water?

motavysium ebboride ammonium nitrate sodium sulfate potassium chromate DO four All twee B) three C) one E) nome

When solutions of silver nitrate and potassium chloride are mixed, which of the following will be present in the net ionic equation?

Balk" A) KNOwe

C) KCl(aq). EDO CIT E) None of these

3. The net ionic equation for formation of an aqueous solution of Nil, accompanied by evolution of COgas via mixing solid NiCO, and aqueous hydroiodic acid is _

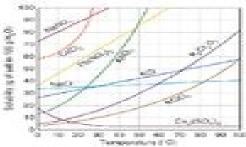
 A_1 2NiCO (s) + HI(aq) \rightarrow 2H (O(1) + CO (g) + 2Ni (aq)

P(s) = P(s) + P(s) +

 $O(NiCO_1(s) + 2H^*(aq) \rightarrow H_1O(l) + CO_2(g) + Ni^{**}(aq)$

Da NiCO (is) + 2HI(aq) - 2H (O(I) + CO (g) + NiI (aq)

 $P_1 \text{NiCO}_1(s) + 2HI(aq) - H_1O(l) + CO_1(g) + Ni^{-1}(aq) + 2I_1(aq)$



 A beaker containing 80 grams of lead(II) nitrate, Pb(NO₁)_c, in 100 grams of water has a temperature of 30 °C. Approximately how many grams of the salt are undissolved, on the bottom of the beaker? A) 14 grams B) 66 grams. (C) 80 grams. D) 20 grams.

 A solution in which no more solid can dissolve is called a sublittion:

A) solute: F) soluble B) solution G5 solvient:

C) electrolyte D) immiscible E) concentrated.

IK) saturated.

LA moisseible:

H) unsaturated (I) dilute

J) non-electrolyte

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