

FIITJEE

CHEMICAL EQUILIBRIUM

EXERCISE – I

OBJECTIVE PROBLEMS (JEE MAIN)

1. For the reaction equilibrium :

$\text{N}_2\text{O}_4(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$; the concentration of N_2O_4 and NO_2 at equilibrium are 4.8×10^{-2} and 1.2×10^{-2} mol/L respectively. The value of K_c for the reaction is :

- (A) 3×10^{-2} M (B) 3×10^2 M (C) 3.3×10^2 M (D) 3×10^{-4} M

2. What is the equilibrium constant for the reaction
- $\text{P}_{40(\text{s})} + 5\text{O}_{2(\text{g})} \rightleftharpoons \text{P}_4\text{O}_{10(\text{s})}$
- :

- (A) $K_c = \frac{1}{[\text{O}_2]^5}$ (B) $K_c = \frac{[\text{P}_4\text{O}_{10}]}{5[\text{P}_4][\text{O}_2]^5}$ (C) $K_c = [\text{O}_2]^5$ (D) $K_c = \frac{[\text{P}_4\text{O}_{10}]}{[\text{P}_4][\text{O}_2]^5}$

3. The equilibrium constant for the reaction :

$\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}(\text{g})$ at temperature

T is 4×10^{-4} . The value of K_c for the reaction.

$\text{NO}(\text{g}) \rightleftharpoons \frac{1}{2}\text{N}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g})$ at the same temperature is :

- (A) 0.02 (B) 50 (C) 4×10^{-4} (D) 2.5×10^{-2}

4. The equilibrium constant for the given reaction :

$\text{SO}_2(\text{g}) \rightleftharpoons \text{SO}_3(\text{g}) + \frac{1}{2}\text{O}_2(\text{g})$; $K_c = 4.9 \times 10^{-2}$

The value of K_c for the reaction :

$2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{SO}_3(\text{g})$, will be

- (A) 416 (B) 2.40×10^{-2} (C) 9.8×10^{-2} (D) 4.9×10^{-2}

5. For the following three reactions 1, 2 and 3, equilibrium constants are given :

(1) $\text{CO}(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightleftharpoons \text{CO}_2(\text{g}) + \text{H}_2(\text{g})$; K_1

(2) $\text{CH}_4(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightleftharpoons \text{CO}(\text{g}) + 3\text{H}_2(\text{g})$; K_2

(3) $\text{CH}_4(\text{g}) + 2\text{H}_2\text{O}(\text{g}) \rightleftharpoons \text{CO}_2(\text{g}) + 4\text{H}_2(\text{g})$; K_3

Which of the following relations is correct ?

- (A) $K_1 \sqrt{K_2} = K_3$ (B) $K_2 K_3 = K_1$ (C) $K_3 = K_1 K_2$ (D) $K_3 = K_2^2 K_1^2$

6. Consider following reactions in equilibrium with equilibrium concentration 0.01M of every species

(I) $\text{PCl}_5(\text{g}) \rightleftharpoons \text{PCl}_3(\text{g}) + \text{Cl}_2(\text{g})$

(II) $2\text{HI}(\text{g}) \rightleftharpoons \text{H}_2(\text{g}) + \text{I}_2(\text{g})$

(III) $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$

Extent of the reactions taking place is :

- (A) $\text{I} > \text{II} > \text{III}$ (B) $\text{I} < \text{II} < \text{III}$ (C) $\text{II} < \text{III} < \text{I}$ (D) $\text{III} < \text{I} < \text{II}$

7. A definite amount of solid
- NH_4HS
- is placed in a flask already containing ammonia gas at a certain temperature and 0.50 atm pressure.
- NH_4HS
- decomposes to give
- NH_3
- and
- H_2S
- and at equilibrium total pressure in flask is 0.84 atm. The equilibrium constant for the reaction is :

- (A) 0.30 (B) 0.18 (C) 0.17 (D) 0.11

Solutions Manual Chemical Equilibrium Chemistry

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