

Ch. 13.2 – Saturated Solutions & Solubility• **Dissolution/ Crystallization Equilibrium:**

- **Dissolution** – as a soluble solid solute dissolves, the number of solute particles in solution _____.
- **Crystallization** – sometimes a dissolved particle will collide with an undissolved portion of the solute and will _____ to the solid.
- The dissolution/crystallization processes are in dynamic _____ which means both processes occur at equal rates (the # of particles dissolving is equal to the # of particles recrystallizing).

• **Types of Solutions:** (these are _____ dependent)○ **Unsaturated Solutions**

- _____ than the maximum amount of solute is dissolved (aqueous)
- at that given temperature
- The solvent has the capability to dissolve _____ solute than what it currently has

○ **Saturated Solutions**

- The solvent has dissolved the _____ amount of solute that it can at that given temperature
- No additional solid can be dissolved in a saturated solution therefore, if any more solid is added, it will remain _____ at the bottom of the beaker

○ **Supersaturated Solution**

- Solvent holds _____ than the maximum amount of solute that can be dissolved at that given temperature
- How is this solution made? ... Heat the solvent, add a TON of _____, and slowly _____ the solution so that as the new, cooler temperature your solution is supersaturated!
- This is an _____ solution; crystallization is easily initiated by adding a "seed crystal" or by agitation/movement

Solutions And Solubility Review Sheet

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