Name ____

Solution Stoichiometry Worksheet

Solve the following solutions Stoichiometry problems:

 How many grams of silver chromate will precipitate when 150, mL of 0,500 M silver nitrate are added to 100, mL of 0,400 M potassium chromate?

2 AgNO_{Net} + K₀CrO_{Net} - Ag₀CrO_{Net} + 2 KNO_{Net}

0.150 L AgNO ₂	0.500 moles AgNO ₂	1 moles Ag ₂ CrO ₄	331.74 g Ag ₂ CrO ₄ = 12.4 g Ag ₂ CrO ₄
	1 L	2 moles AgNO ₃	1 moles Ag ₅ CrO ₄

0.100 L K ₂ CrO ₄	0.400 moles K ₂ CrO ₄	1 moles Ag ₂ CrO ₄	331.74 g Ag ₂ CrO ₄	= 13.3 g AgeCrO.
	1 L	1 moles K ₂ CrO ₄	1 moles Ag ₂ CrO ₄	The state of the s

 How many mL of 0.280 M barium nitrate are required to precipitate as barium sulfate all the sulfate ions from 25.0 mL of 0.350 M aluminum sulfate? (93.8 mL barium nitrate)

3 Ba(NOs)sect + Ale(SOs)sect - 3 BaSOsect + 2 Al(NOs)sect

0.0250 L Al ₂ (SO ₄) ₃	0.350 moles Al ₂ (SO ₄) ₅	3 moles Ba(NO ₃) ₂	1 L	= 0.0938 L. Ba(NO ₃) ₂
	I L	1 moles Al ₂ (SO _a) ₃	0.280 moles Ba(NO ₃) ₂	

 25.0 mL of 0.350 M NaOH are added to 45.0 mL of 0.125 M copper (II) sulfate. How many grams of copper (II) hydroxide will precipitate?

2 NaOH_(ac) + CuSO_{toni} → Cu(OH)_(bc) + Na₀SO_{toni}

0.0250 L NaOH	0.350 moles NaOH	1 moles Cu(OH) ₂	97.57 g Cu(OH) ₂	= 0.427 g Cu(OH) ₂
	1 L NaOH	2 moles NaOH	1 mole Cu(OH) ₂	

0.0450 L CuSO ₄	0.125 moles CuSO,	1 moles Cu(OH) ₂	97.57 g Cu(OH) ₂	= 0.549 g Cu(OH) ₂
	1 L NaOH	1 moles CuSO.	1 mole Cu(OH) ₂	

4. What volume of 0.415 M silver nitrate will be required to precipitate as silver bromide all the bromide ion in 35.0 mL of 0.128 M calcium bromide?

2 AgNO_{3(ac)} + CaBr_{3(ac)} → Ca(NO₃)_{3(ac)} + 2 AgBr_(c)

0.0350 L CaBr ₂	0.128 moles CaBr ₂	2 moles AgNO ₃	1 L AgNO ₃	= 0.0216 L AgNO ₃
	1 L CaBr ₂	1 moles CaBr ₂	0.415 mole AgNO _a	7.5

5. What volume of 0.496 M HCl is required to neutralize 20.0 mL of 0.809 M sodium hydroxide?

HCl(aq) + NaOH(aq) → NaCl(aq) + H(OH)(I)

0.0200 L NaOH	0.809 mole NaOH	1 mole HCI	1 L HCI	= 0.0326 L HCI
	1 L NaOH	1 mole NaOH	0.496 mole HCI	

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Table of Contents Solution Stoichiometry Name Chem Worksheet 146876

- 1. Understanding the eBook Solution Stoichiometry Name Chem Worksheet 146876
 - The Rise of Digital Reading Solution Stoichiometry Name Chem Worksheet 146876
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Solution Stoichiometry Name Chem Worksheet 146876
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solution Stoichiometry Name Chem Worksheet 146876
 - User-Friendly Interface

- 4. Exploring eBook Recommendations from Solution Stoichiometry Name Chem Worksheet 146876
 - Personalized Recommendations
 - Solution Stoichiometry Name Chem Worksheet 146876 User Reviews and Ratings
 - Solution Stoichiometry Name Chem Worksheet 146876 and Bestseller Lists
- 5. Accessing Solution Stoichiometry Name Chem Worksheet 146876 Free and Paid eBooks
 - Solution Stoichiometry Name Chem Worksheet 146876 Public Domain eBooks
 - Solution Stoichiometry Name Chem Worksheet 146876 eBook Subscription Services
 - Solution Stoichiometry Name Chem Worksheet 146876 Budget-Friendly Options
- 6. Navigating Solution Stoichiometry Name Chem Worksheet 146876 eBook Formats
 - ∘ ePub, PDF, MOBI, and More
 - Solution Stoichiometry Name Chem Worksheet 146876 Compatibility with Devices
 - Solution Stoichiometry Name Chem Worksheet 146876 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Solution Stoichiometry Name Chem Worksheet 146876
 - Highlighting and Note-Taking Solution Stoichiometry Name Chem Worksheet 146876
 - Interactive Elements Solution Stoichiometry Name Chem Worksheet 146876
- 8. Staying Engaged with Solution Stoichiometry Name Chem Worksheet 146876
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solution Stoichiometry Name Chem Worksheet 146876
- 9. Balancing eBooks and Physical Books Solution Stoichiometry Name Chem Worksheet 146876
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solution Stoichiometry Name Chem Worksheet 146876
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Solution Stoichiometry Name Chem Worksheet 146876
 - Setting Reading Goals Solution Stoichiometry Name Chem Worksheet 146876
 - Carving Out Dedicated Reading Time

- 12. Sourcing Reliable Information of Solution Stoichiometry Name Chem Worksheet 146876
 - Fact-Checking eBook Content of Solution Stoichiometry Name Chem Worksheet 146876
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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