

Chem 213-219
Date: 10/19/2020

Experiment 007: The Titration of Vinegar

References:

Suzanne Slayden, 2018, Chemistry 213, 214, 272 Laboratory Experiments, 7th Edition, pages 19-23.

George Mason University, Department of Chemistry and Biochemistry. Name of the Lab Video; 2020.

The Titration of Vinegar, Distance Learning Fall 2020, accessed via Blackboard.

Purpose:

A common household item, vinegar, contains acetic-acid in addition to other chemicals. The purpose of this experiment is to identify the amount of acetic acid present in vinegar, which is essentially an acetic-acid solution. This will be accomplished by titrating the solution with a standard base, Sodium hydroxide. The anticipated outcome of this experiment is that as the base, Sodium hydroxide, is added to the acetic-acid, the result will yield a neutralization reaction. Additionally, the amount of moles processed to neutralize the acetic-acid solution in Sodium hydroxide must be the same number of moles determined in vinegar.

Materials:

- 10-mL of vinegar
- Sodium hydroxide solution
- Oxalic acid
- 10-mL graduated cylinder
- Electronic balance
- Distilled water
- Phenolphthalein indicator
- 250-mL erlenmeyer flask
- Burrett
- Retort stand
- Funnel
- Pipette
- 100-mL volumetric flask
- White piece of paper (this will help distinguish the color change throughout the experiment)
- Gloves
- Goggles

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