

Writing Formulas and Naming Compounds

Directions: Answer the following questions in the spaces provided. Refer to the periodic table for help.

1. Define an oxidation number. Is the number of electrons an atom will **lose** **gain** to become a positive ion or **lose** **gain** to become a negative ion and have a full outer energy level of electrons
2. What is the usual oxidation number of oxygen? Of hydrogen? O H
3. What is the sum of all the oxidation numbers in any compound? _____ use a number

Directions: Use the periodic table to identify the oxidation numbers of the elements in each group.

Group	1	2	6	7	8
Oxidation number					

Directions: Write the formulas for the following compounds. Use the periodic table for help.

Symbols with oxidation numbers

Formula

10. copper(II) sulfate _____
11. calcium chloride _____
12. iron(II) oxide _____
13. copper(I) oxide _____
14. sodium sulfide _____
15. magnesium bromide _____

Directions: complete the following table, providing the name of the compound and the total number of atoms in each formula given.

Formula	Name	Number of Atoms
16. $(\text{NH}_4)\text{OH}$		
17. $(\text{NH}_4)\text{Cl}$		
18. Ag_2O		
19. $\text{K}_2(\text{SO}_4)$		
20. $\text{Ca}(\text{NO}_3)_2$		
21. Na_2S		

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Answer Key

Virginia P. Powell



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