

Answer KeyNaming and Writing Formulas for Ionic Compounds

Write the names for the following compounds:

1. KClO_4 – potassium perchlorate
2. Sr_3P_2 – strontium phosphide
3. NH_4NO_2 – ammonium nitrite
4. NaCl – sodium chloride
5. CoS – cobalt (II) sulfide
6. ZnCl_2 – zinc (II) chloride
7. Li_3N – lithium nitride
8. BaSO_3 – barium sulfite
9. Co_2S_3 – cobalt (III) sulfide
10. CsF – cesium fluoride
11. NaNO_3 – sodium nitrate
12. CuI_2 – copper (II) iodide
13. TiO_2 – titanium (IV) oxide
14. KMnO_4 – potassium permanganate
15. ZnCN_2 – zinc (II) cyanide
16. $\text{Al}_2(\text{CO}_3)_3$ – aluminum carbonate
17. $\text{Fe}_2(\text{CO}_3)_3$ – iron (III) carbonate
18. BaCrO_4 – barium chromate
19. $\text{Pb}_3(\text{PO}_4)_2$ – lead (II) phosphate
20. NaHCO_3 – sodium hydrogen carbonate or sodium bicarbonate

Write the formulas for the following compounds:

1. cesium bromide – CsBr
2. aluminum chlorite – $\text{Al}(\text{ClO}_2)_3$
3. ammonium phosphite – $(\text{NH}_4)_3\text{PO}_3$
4. copper (II) nitrate – $\text{Cu}(\text{NO}_3)_2$
5. sodium fluoride – NaF
6. lead (IV) sulfate – $\text{Pb}(\text{SO}_4)_2$
7. calcium carbonate – CaCO_3
8. barium oxide – BaO
9. iron (III) sulfide – Fe_2S_3
10. manganese (II) chromate – MnCrO_4

Writing Ionic Formulas Lab 2answer Key

Gary S. Thorpe

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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