Name:		

Writing and Balancing Equations — Alternate Option

Procedure:

For each chemical word equation below, first write the formula for each reactant and product. Then write the complete chemical equation; include the physical state for each substance. Balance each equation by following the procedure outlined on page 3 of the handout.

- Hydrogen and oxygen gas react to form water.
- Aqueous lead(II) chloride reacts with aqueous sodium sulfate to produce a lead(II) sulfate precipitate and aqueous sodium chloride.
- 3. Silicon reacts with chlorine gas to form liquid silicon tetrachloride.
- 4. Aqueous chlorine and aqueous potassium bromide react to make a solution of potassium chloride and bromine.
- Sodium metal combines with chlorine gas to form crystalline sodium chloride.
- Nitrogen combines with hydrogen to make gaseous ammonia.
- Solid sodium bicarbonate decomposes to make aqueous sodium carbonate, carbon dioxide gas, and water.
- Aluminum combines with oxygen gas to form solid aluminum oxide.
- 9. Methane gas undergoes combustion with oxygen to produce carbon dioxide and water.
- Zinc reacts with aqueous hydrogen chloride to form aqueous zinc chloride and hydrogen gas.
- Lithium metal combines with oxygen to form lithium oxide.
- 12. Aqueous silver nitrate reacts with solid copper to form a cupric nitrate solution and silver crystals.
- Propane fuel burns in oxygen to form carbon dioxide and water.
- 14. Solid calcium carbonate decomposes to produce calcium oxide and carbon dioxide.
- 15. A sodium hydroxide solution neutralizes a hydrochloric acid solution to form a solution of sodium chloride and water.
- 16. Solid mercuric oxide decomposes to form liquid mercury metal and oxygen.
- 17. Iron metal reacts with water to form iron(III) oxide and hydrogen gas.
- 18. Aluminum reacts with a solution of hydrochloric acid to form aluminum chloride dissolved in water.
- 19. Potassium metal reacts violently with water to form a basic solution of potassium hydroxide, releasing hydrogen gas.
- 20. Ammonia gas and oxygen combine to form nitrous oxide and water vapor.

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le domaine est ouvert de 12h à 18h30 en haute saison d'avril à octobre et de 12h à 17h30 en basse saison de novembre à mars en juillet et août l'ouverture a lieu exceptionnellement à 10h du mardi au dimanche les jardins restent ouverts jusqu à 19h30 en haute saison et jusqu à 18h en basse saison

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au fond du domaine le long des rives du lac se dressent les chaumières du hameau construit par richard mique pour marie antoinette survivantes d une époque brillante et mouvementée elles illustrent le goût de la reine pour le charme de la vie champêtre

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located in an 18th century farmhouse and offering a view of the valley and the village of saint paul de vence le hameau features an outdoor swimming pool and relaxation area free wifi is available throughout the property