

# CHEMICAL BONDING

↳ Force of attraction present between any chemical species like atoms, ions & molecules.

Bond formation → Exothermic

Bond dissociation → Endothermic

## Chemical bonds on the basis of Bond Energy

### Strong bond

- ① Ionic bond
- ② Covalent bond
- ③ co-ordinate bond
- ④ Metallic bond

### Weak bond

- ① Hydrogen bond
- ② Vander waal's bond

## # Cause of chemical combination

# Modern concept :- Every atom has a tendency to form max<sup>m</sup> possible bonds in order to minimize its potential energy & gain stability.

## # Classical concept (octet rule) :-

↳ Every atom has a tendency to achieve stable outer octet ( $2e^-$  in valence shell).

## # Different theories to explain bonding

- ① Octet rule
- ② VSEPR (valence shell electron pair repulsion theory)
- ③ VBT (valence bond theory)
- ④ MOT (Molecular orbital theory)

## # Kosser - Lewis Approach (octet rule) :-

- Based on inertness of noble gas
- Atoms can be combine either by transfer or by sharing of valence electrons in order to have an octet ( $2e^-$  in their valence shells).
- In case of H & He duplet ( $2e^-$  in valence shell) will be completed.

## # How to write Lewis dot structure & bond formation of molecules

$$\text{No. of Bonds} = \frac{(\text{No. of atoms} \times \text{Req. valence } e^-) - \text{Available val. } e^-}{2}$$



$$\text{No. of bonds} = \frac{(2 \times 2) - (4 + 6)}{2} = 3$$

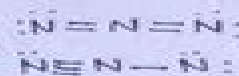
btw C & O



$$\text{No. of bonds} = \frac{(8 \times 3) - (6 \times 3)}{2} = 3$$



$$\text{No. of bonds} = \frac{8 \times 2 - (5 \times 3) + 1}{2} = 4$$



## # Limitations of octet rule

(i) Existence of Hypervalent compound.



(ii) Existence of Hypovalent compound.



(iii) Odd electron species do not follow octet rule. e.g. → NO, NO2

(iv) Existence of compounds of inert gases with oxygen & fluorine like XeF2, KrF2, XeF4 etc.

(v) does not explain the relative stability of the molecules being totally silent about the energy of a molecule.

(vi) This theory does not account for the shape of molecules.

## ⇒ formal charge :-

F.C = Jitne hone chahiye - Jitne hain on atom (val.  $e^-$  in free state) (val.  $e^-$  in given structure)

$$\text{F.C} = \text{Total val. } e^- - [2 \times \text{L.P} + \text{Bond pair}]$$

# Short Guide To Writing About Chemistry

**B Lingard**



## **Short Guide To Writing About Chemistry:**

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### **Table of Contents Short Guide To Writing About Chemistry**

1. Understanding the eBook Short Guide To Writing About Chemistry
  - The Rise of Digital Reading Short Guide To Writing About Chemistry
  - Advantages of eBooks Over Traditional Books
2. Identifying Short Guide To Writing About Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Short Guide To Writing About Chemistry
  - User-Friendly Interface
4. Exploring eBook Recommendations from Short Guide To Writing About Chemistry
  - Personalized Recommendations

- Short Guide To Writing About Chemistry User Reviews and Ratings
- Short Guide To Writing About Chemistry and Bestseller Lists
- 5. Accessing Short Guide To Writing About Chemistry Free and Paid eBooks
  - Short Guide To Writing About Chemistry Public Domain eBooks
  - Short Guide To Writing About Chemistry eBook Subscription Services
  - Short Guide To Writing About Chemistry Budget-Friendly Options
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  - Short Guide To Writing About Chemistry Enhanced eBook Features
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  - Highlighting and Note-Taking Short Guide To Writing About Chemistry
  - Interactive Elements Short Guide To Writing About Chemistry
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  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Short Guide To Writing About Chemistry
- 9. Balancing eBooks and Physical Books Short Guide To Writing About Chemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Short Guide To Writing About Chemistry
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Short Guide To Writing About Chemistry
  - Setting Reading Goals Short Guide To Writing About Chemistry
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Short Guide To Writing About Chemistry
  - Fact-Checking eBook Content of Short Guide To Writing About Chemistry

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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