

Chemical Formula Writing Worksheet Solutions

Write chemical formulas for the compounds in each box. The names are found by finding the intersection between the cations and anions. Example: The first box is the intersection between the "zinc" cation and the "chloride" anion, so you should write " ZnCl_2 ", as shown.

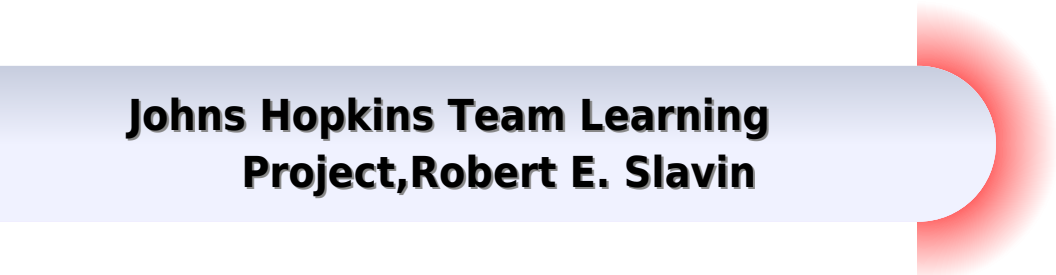
	<i>zinc</i>	<i>iron (II)</i>	<i>iron (III)</i>	<i>gallium</i>	<i>silver</i>	<i>lead (IV)</i>
<i>chloride</i>	ZnCl_2	FeCl_2	FeCl_3	GaCl_3	AgCl	PbCl_4
<i>acetate</i>	$\text{Zn}(\text{C}_2\text{H}_3\text{O}_2)_2$	$\text{Fe}(\text{C}_2\text{H}_3\text{O}_2)_2$	$\text{Fe}(\text{C}_2\text{H}_3\text{O}_2)_3$	$\text{Ga}(\text{C}_2\text{H}_3\text{O}_2)_3$	$\text{Ag C}_2\text{H}_3\text{O}_2$	$\text{Pb}(\text{C}_2\text{H}_3\text{O}_2)_4$
<i>nitrate</i>	$\text{Zn}(\text{NO}_3)_2$	$\text{Fe}(\text{NO}_3)_2$	$\text{Fe}(\text{NO}_3)_3$	$\text{Ga}(\text{NO}_3)_3$	AgNO_3	$\text{Pb}(\text{NO}_3)_4$
<i>oxide</i>	ZnO	FeO	Fe_2O_3	Ga_2O_3	Ag_2O	PbO_2
<i>nitride</i>	Zn_3N_2	Fe_3N_2	FeN	GaN	Ag_3N	Pb_3N_4
<i>sulfate</i>	ZnSO_4	FeSO_4	$\text{Fe}_2(\text{SO}_4)_3$	$\text{Ga}_2(\text{SO}_4)_3$	Ag_2SO_4	$\text{Pb}(\text{SO}_4)_2$

Write the formulas for the following compounds:

- 1) copper (II) chloride **CuCl_2**
- 2) lithium acetate **$\text{LiC}_2\text{H}_3\text{O}_2$**
- 3) vanadium (III) selenide **V_2Se_3**
- 4) manganese (IV) nitride **Mn_3N_4**
- 5) beryllium oxide **BeO**
- 6) sodium sulfate **Na_2SO_4**
- 7) aluminum arsenide **AlAs**
- 8) potassium permanganate **KMnO_4**
- 9) chromium (VI) cyanide **$\text{Cr}(\text{CN})_6$**
- 10) tin (II) sulfite **SnSO_3**
- 11) vanadium (V) fluoride **VF_5**
- 12) ammonium nitrate **NH_4NO_3**

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**Johns Hopkins Team Learning
Project, Robert E. Slavin**



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