

UNEB UACE CHEMISTRY PAPER 3 2018

1. You are provided with the following:

FA1, which is approximately a 0.1M sodium thiosulphate solution

FA2, which is a solution containing 2.4g dm⁻³ of potassium iodate

Solid Y, which is a salt containing dichromate ions

1M sulphuric acid solution

5% potassium iodide solution

Starch solution.

You are required to standardize FA1 and use it to determine the percentage by mass of chromonium in Y.

In acidic solution, iodate and dichromate (VI) ions react with potassium iodide to liberate iodine according to the following equations.



The iodine liberated in both cases reacts with thiosulphate ions according to the following equation



PROCEDURE

a) Pipette 25.0 (or 20.0)cm³ of FA2 into a conical flask and add an equal volume of 1M sulphuric acid using a measuring cylinder, followed by 10cm³ of 5% potassium iodide solution. Titrate the iodine liberated with FA1, using starch solution as indicator. Repeat the titration until you obtain consistent results.

i) Record your results in Table I below.

Results:

Volume of pipette used.....cm³

Table I

Final burette reading (cm ³)			
Initial burette reading (cm ³)			
Volume of FA1 used (cm ³)			

ii) Volumes of FA1 used for calculating average volume.....cm³

iii) Average volume of FA1 used.....cm³

Questions:

a) Calculate the number of moles of iodine liberated by FA2 (O=16; K=39; I= 127)

b) Determine the concentration of FA1 in mol dm⁻³

PROCEDURE:

b) Weigh accurately about 1.2g of Y. Dissolve it in a minimum amount of distilled water and transfer the solution into a 250cm³ volumetric flask. Make the solution up to the mark with distilled water and label it FA3.

Pipette 25.0 (or 20.0) cm³ of FA3 into a conical flask and add an equal volume of 1M sulphuric acid using a measuring cylinder, followed by 10cm³ of 5% potassium iodide solution.

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