

Moles to Particles Conversions

Name _____

Complete the following conversions. For FULL CREDIT, you must show ALL of your work!

1. Determine the number of atoms in 0.58 moles of Se.

$$0.58 \text{ mol Se} \times \frac{6.02 \times 10^{23} \text{ atoms Se}}{1 \text{ mol Se}} = 3.5 \times 10^{23} \text{ atoms Se}$$

2. How many moles are in 6.80×10^{24} formula units $\text{Ba}(\text{NO}_3)_2$?

$$6.80 \times 10^{24} \text{ formula units Ba(NO}_3)_2 \times \frac{1 \text{ mol Ba(NO}_3)_2}{6.02 \times 10^{23} \text{ formula units Ba(NO}_3)_2} = 11.3 \text{ mol Ba(NO}_3)_2$$

3. Determine the number of molecules in 1.25 moles of O_2 gas.

$$1.25 \text{ mol O}_2 \times \frac{6.02 \times 10^{23} \text{ molecules O}_2}{1 \text{ mol O}_2} = 7.52 \times 10^{23} \text{ molecules O}_2$$

4. How many moles are in 5.38×10^{24} formula units MgBr_2 ?

$$5.38 \times 10^{24} \text{ formula units MgBr}_2 \times \frac{1 \text{ mol MgBr}_2}{6.02 \times 10^{23} \text{ formula units MgBr}_2} = 8.94 \text{ mol MgBr}_2$$

5. Determine the number of formula units in 0.688 moles of AgNO_3 .

$$0.688 \text{ mol AgNO}_3 \times \frac{6.02 \times 10^{23} \text{ formula units AgNO}_3}{1 \text{ mol AgNO}_3} = 4.14 \times 10^{23} \text{ formula units AgNO}_3$$

6. Convert 8.46×10^{24} molecules of ethane (C_2H_6) to moles.

$$8.46 \times 10^{24} \text{ molecules C}_2\text{H}_6 \times \frac{1 \text{ mol C}_2\text{H}_6}{6.02 \times 10^{23} \text{ molecules C}_2\text{H}_6} = 14.1 \text{ mol C}_2\text{H}_6$$

7. Determine the number of formula units in 1.48 moles of NaF .

$$1.48 \text{ mol NaF} \times \frac{6.02 \times 10^{23} \text{ formula units NaF}}{1 \text{ mol NaF}} = 8.91 \times 10^{23} \text{ formula units NaF}$$

8. Convert 8.2×10^{25} molecules of water to moles.

$$8.2 \times 10^{25} \text{ molecules H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{6.02 \times 10^{23} \text{ molecules H}_2\text{O}} = 140 \text{ mol H}_2\text{O}$$

9. How many formula units are in 15 moles of Li_2SO_4 ?

$$15 \text{ mol Li}_2\text{SO}_4 \times \frac{6.02 \times 10^{23} \text{ formula units Li}_2\text{SO}_4}{1 \text{ mol Li}_2\text{SO}_4} = 9.0 \times 10^{24} \text{ formula units Li}_2\text{SO}_4$$

10. Calculate the number of moles in 7.2×10^{23} atoms of sulfur.

$$7.2 \times 10^{23} \text{ atoms S} \times \frac{1 \text{ mol S}}{6.02 \times 10^{23} \text{ atoms S}} = 0.12 \text{ mol S}$$

*Answer rounded using significant figures even/odd rounding rules.

Skills Practice 31 Moles And Formulas

R Barnett



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Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

The Practice of Chemistry Donald J. Wink, Sharon Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book's unique focus on actual chemical practice extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink http://www.bfwpub.com/pdfs/wink/POCPowerPoint_Final.ppt 832KB

The Software Finder, 1983 **Prentice Hall Chemistry**, 2000 A-Level Chemistry for AQA: Year 1 & 2 Student Book, 2020-10-05 This fantastic CGP Student Book comprehensively covers both years of AQA A Level Chemistry It's bursting with in depth accessible notes explaining every course topic plus all of the Required Practicals Everything's supported by clear diagrams photographs tips and worked examples Throughout the book there are lots of practice questions and exam style questions with answers at the back There's detailed guidance on Maths Skills and Practical Skills as well as indispensable advice for success in the final exams If you'd prefer Year 1 9781782943211 Year 2 9781782943266 in separate books CGP has them too And for more detailed coverage of the mathematical elements of A Level Chemistry try our Essential Maths Skills book 978182944720

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Glencross, Qiuyu Wang, 2016 Biomedical scientists are the foundation of modern healthcare from cancer screening to diagnosing HIV from blood transfusion for surgery to food poisoning and infection control Without biomedical scientists the diagnosis of disease the evaluation of the effectiveness of treatment and research into the causes and cures of disease would not be possible The Fundamentals of Biomedical Science series has been written to reflect the challenges of practicing biomedical science today It draws together essential basic science with insights into laboratory practice to show how an understanding of the biology of disease is coupled to the analytical approaches that lead to diagnosis Assuming only a minimum of prior knowledge the series reviews the full range of disciplines to which a Biomedical Scientist may be exposed from microbiology to cytopathology to transfusion science A core text in the Fundamentals of Biomedical Science series Biomedical Science Practice gives a comprehensive overview of the key laboratory techniques and professional skills that students need to master The text is supported throughout with engaging clinical case studies written to emphasize the link between theory and practice providing a strong foundation for beginning biomedical science students Ebook: Chemistry

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