

Mole Conversion Practice

Name: _____

Date: _____

Hour: _____

1. Calculate the number of molecules in 210 grams of water.

7.02x10²⁴ molecules

2. If you have 6.25×10^{24} molecules of sulfur tetrafluoride, how many grams do you have?

1120 g

3. Define what a mole is. Don't give a number, give a definition in words.

The quantity of atoms needed to equal the substance's atomic mass in units of grams.

4. What is the mass of 3.45 moles of aluminum sulfate?

1180 g

5. Consider zinc acetate and find its "formula mass" and its "molar mass". Include units for each.

Formula mass: 183.4 amu Molar mass: 183.4 g/mol

6. If you have 245 grams of lithium carbonate, how many moles do you have?

3.32 mol

7. How many gold atoms do you have if you have 400 grams of gold?

1.22x10²⁴ atoms

8. 68.5 grams of a certain compound contains 5.25×10^{23} molecules. What is the molar mass of the compound?

78.5 g/mol

9. How many grams are there in a container of 2.26×10^{25} molecules of dinitrogen trisulfide?

4,670 g

10. How many molecules are in 325 g of N₂O₃?

1.91x10²⁴ molecules

Skill Practice 30 Chemistry

JS Bruner



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