

DIAGNOSTIC TEST

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Time: 60 minutes

Part A

Note: Unless otherwise stated, for all statements involving chemical equations and/or solutions, assume that the system is in pure water.

Directions: Each of the following sets of lettered choices refers to the numbered formulas or statements immediately below it. For each numbered item, choose the one lettered choice that fits it best. Then fill in the corresponding oval on the answer sheet. Each choice in a set may be used once, more than once, or not at all.

Questions 1–4

- (A) $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$
- (B) $\text{CH}_3\text{CH}(\text{Br})\text{CH}(\text{Br})\text{CH}_3$
- (C) $\text{CH}_3\text{CH}_2\text{CH}(\text{Br})\text{CH}_3$
- (D) $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}$
- (E) $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$



(3)

4. Which compound would most likely turn litmus paper to a red color?

Questions 5–8

- (A) Heisenberg Uncertainty Principle
- (B) Pauli Exclusion Principle
- (C) Schrödinger Wave Equation
- (D) Hund's rule
- (E) Bohr model of the hydrogen atom

- 5. No two electrons can have the same quantum number because they must have opposite spins.
- 6. We cannot know the exact location of an electron in space.
- 7. The electrons will occupy an orbital singly, with parallel spins, before pairing up.

8. The energy changes that an electron may undergo are quantized.

Questions 9–12

- (A) H_2
- (B) CO_2
- (C) H_2O
- (D) NaCl
- (E) CH_3CH_3

- 9. Contains just one sigma bond.
- 10. Has a bond formed from the transfer of electrons.
- 11. Has an atom that is sp hybridized.
- 12. Is a polar molecule.

Questions 13–16

- (A) F
- (B) Li
- (C) Fe
- (D) He
- (E) Si

- 13. Shows both the properties of both metals and non-metals.
- 14. Has the greatest ionization energy.
- 15. Has the greatest electronegativity.
- 16. Has colored salts that will produce colored aqueous solutions.

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