#### CHAPTER 4 PRACTICE TEST: SOLUTION STOICHIOMETRY

1) Which of the following are weak electrolytes? HCl, HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>, NH<sub>3</sub>, KCl
A) HCl, KCl
B) HCl, HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>, NH<sub>3</sub>, KCl
C) HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>, KCl
Answer: D Diff: 2

2) What are the spectator ions in the reaction between KOH (aq) and HNO, (aq)?

A) K'and H' C) K'and NO<sub>3</sub> E) OH only B) H'and OH D) H'and NO<sub>3</sub> Answer: C Diff: 2

3) The balanced net ionic equation for precipitation of CaCO, when aqueous solutions of

Na 2 CO 3 and CaCl 2 are mixed is \_\_\_\_\_\_.

A)  $2Na^{\circ}(aq) + CO_3^{\circ}(aq) \rightarrow Na_2 CO_3(aq)$  E)

B)  $2 \text{ Na}^*(aq) + 2\text{Cl}^*(aq) \rightarrow 2\text{NaCl}(aq)$  Na  $_2\text{CO}_3(aq) + \text{CaCl}_2(aq) \rightarrow 2\text{NaCl}(aq) + \text{CaCO}_3(s)$ 

C) Na\*(aq)+Cl\*(aq) → NaCl (aq)

D)  $Ca^{2+}(aq) + CO_3^{2-}(aq) \rightarrow CaCO_3(s)$ 

Answer: D Diff: 3

4) Which of the following are strong acids? HI, HNO<sub>3</sub>, HF, HBr

A) HF, HBr C) HI, HF, HBr

B) HI, HNO, HF, HBr D) HNO, HF, HBr

Answer: E Diff: 2

5) A neutralization reaction between an acid and a metal hydroxide produces

A) water and a salt C) oxygen gas E) ammonia

B) hydrogen gas D) sodium hydroxide

Answer: A Diff: 2

6) Which of these metals is the least easily oxidized? Na, Au, Fe, Ca, Ag

A) Na B) Au C) Fe D) Ca

Answer: B Diff: 3

7) Based on the equations below, which metal is the most active?

 $Pb(NO_3)_2(aq) + Ni(s) \rightarrow Ni(NO_3)_2(aq) + Pb(s)$ 

 $Pb(NO_1)_1(aq) + Ag(s) \rightarrow No reaction$ 

 $Cu(NO_3)_3(aq) + Ni(s) \rightarrow Ni(NO_3)_3(aq) + Cu(s)$ 

A) Ni B) Ag C) Cu Answer: A Diff: 3

8) What is the concentration (M) of KCl in a solution made by mixing 25.0 mL of 0.100 M KCl with 50.0 mL of 0.100 M KCl?

A) 0.100 B) 0.0500 Answer: A Diff: 3 C) 0.0333

D) 0.0250

D) Pb

E) 125

E) Ag

E) N

E) HI, HNO, HBr

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