

Mock Acid-Base Calculation Test

NOTE TO ASSESSORS:

The questions below should not be used as the "Acid-Base Calculations Test" which is part of Unit 13: Applications of Inorganic Chemistry - Assignment 'Acid or Base?'. It may however be used as a guide to produce your centre's version of the "Acid-Base Calculations Test".

NOTE TO LEARNERS:

Complete the following Mock Test in preparation for the "Acid-Base Calculations Test" which is part of Unit 13: Applications of Inorganic Chemistry - Assignment 'Acid or Base?'

Table of dissociation constant data

Weak acid	$K_a / \text{mol dm}^{-3} \text{ at } 25^\circ\text{C}$
Ethanoic acid	1.7×10^{-5}
Methanoic acid	1.8×10^{-4}
Phosphonic acid	7.1×10^{-3}
Tartaric Acid	9.2×10^{-4}

Formulae for questions:

To calculate pH from concentration of hydrogen ions, $[\text{H}^+]$, and vice-versa

$$\text{pH} = -\log [\text{H}^+] \quad \text{and} \quad [\text{H}^+] = 10^{-\text{pH}}$$

For strong monoprotic acids of formula HA, assume that dissociation into ions is complete, so

$$[\text{H}^+] = [\text{HA}]$$

For strong diprotic acids of formula $[\text{H}_2\text{A}]$,

$$[\text{H}^+] = 2 \times [\text{H}_2\text{A}]$$

For weak acids of formula HA,

For alkalis, concentration of H^+ ions is

where the ionic dissociation constant of water, $K_w = 1 \times 10^{-14} \text{ mol}^2 \text{ dm}^{-6} \text{ at } 25^\circ\text{C}$

For strong alkalis of formula MOH, assume that dissociation into ions is complete, so

$$[\text{OH}^-] = [\text{MOH}]$$

For strong alkalis of formula $\text{M}(\text{OH})_2$,

$$[\text{OH}^-] = 2 \times [\text{M}(\text{OH})_2]$$

For buffer solutions, where $[\text{HA}]$ = concentration of weak and $[\text{A}^-]$ = concentration of conjugate base,

(This is the Henderson-Hasselbach equation)

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Table of Contents Tesccc Unit 13 Lesson One Chemistry

1. Understanding the eBook Tesccc Unit 13 Lesson One Chemistry
 - The Rise of Digital Reading Tesccc Unit 13 Lesson One Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Tesccc Unit 13 Lesson One Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Tesccc Unit 13 Lesson One Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Tesccc Unit 13 Lesson One Chemistry
 - Personalized Recommendations
 - Tesccc Unit 13 Lesson One Chemistry User Reviews and Ratings
 - Tesccc Unit 13 Lesson One Chemistry and Bestseller Lists
5. Accessing Tesccc Unit 13 Lesson One Chemistry Free and Paid eBooks
 - Tesccc Unit 13 Lesson One Chemistry Public Domain eBooks
 - Tesccc Unit 13 Lesson One Chemistry eBook Subscription Services

- Tesccc Unit 13 Lesson One Chemistry Budget-Friendly Options
- 6. Navigating Tesccc Unit 13 Lesson One Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Tesccc Unit 13 Lesson One Chemistry Compatibility with Devices
 - Tesccc Unit 13 Lesson One Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Tesccc Unit 13 Lesson One Chemistry
 - Highlighting and Note-Taking Tesccc Unit 13 Lesson One Chemistry
 - Interactive Elements Tesccc Unit 13 Lesson One Chemistry
- 8. Staying Engaged with Tesccc Unit 13 Lesson One Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Tesccc Unit 13 Lesson One Chemistry
- 9. Balancing eBooks and Physical Books Tesccc Unit 13 Lesson One Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Tesccc Unit 13 Lesson One Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Tesccc Unit 13 Lesson One Chemistry
 - Setting Reading Goals Tesccc Unit 13 Lesson One Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Tesccc Unit 13 Lesson One Chemistry
 - Fact-Checking eBook Content of Tesccc Unit 13 Lesson One Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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