

Atomic Emission Spectra Lab

Directions: From the lab video, record the color of flame produced by each of the metals. Use the chart provided to find the approximate wavelength of the light of each flame using the chart.

1. Before beginning, record the atomic number of the metal in each salt listed below (the metal is the first element listed).
2. Use the equation $c = \lambda \cdot \nu$ (speed of light = frequency x wavelength) to calculate the frequency. Be sure to convert the wavelength of meters before you put it into the equation. ($1 \text{ m} = 1 \times 10^9 \text{ nm}$)
3. Use the frequency in the equation $E = h \cdot \nu$ ($h = 6.63 \times 10^{-34} \text{ J}\cdot\text{s}$) to calculate the energy of the light.

**** You must show your work on all calculations to receive credit! Take a picture of your work and upload it to Canvas with your Lab. You can submit more than one document. ***

Color	Wavelength (nanometers)
Red Light	~650 nm
Orange Light	~590 nm
Yellow Light	~570 nm
Green Light	~510 nm
Blue Light	~475 nm
Indigo Light	~445 nm
Violet Light	~400 nm



Metal Ion Tested	Atomic Number	Color of Flame	Wavelength (m)	Frequency (Hz)	Energy (J)
KCl	19	violet	$4.00 \times 10^{-7} \text{ m}$	$7.5 \times 10^{14} \text{ Hz}$	4.97×10^{-19}
CuCl ₂	29	green	$5.10 \times 10^{-7} \text{ m}$	$5.9 \times 10^{14} \text{ Hz}$	3.91×10^{-19}
CaCl ₂	20	orange	$5.90 \times 10^{-7} \text{ m}$	$5.1 \times 10^{14} \text{ Hz}$	3.38×10^{-19}
NaCl	11	yellow	$5.70 \times 10^{-7} \text{ m}$	$5.3 \times 10^{14} \text{ Hz}$	3.51×10^{-19}
SrCl ₂	38	red	$6.50 \times 10^{-7} \text{ m}$	$4.6 \times 10^{14} \text{ Hz}$	3.05×10^{-19}
LiCl	3	red	$6.50 \times 10^{-7} \text{ m}$	$4.6 \times 10^{14} \text{ Hz}$	3.05×10^{-19}

Part 2- Describe the incandescent and fluorescent lights and how they look different through the spectrum filter.

Virtual Chem Lab 5 Atomic Emission Spectra

Lingsheng Yao



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Table of Contents Virtual Chem Lab 5 Atomic Emission Spectra

1. Understanding the eBook Virtual Chem Lab 5 Atomic Emission Spectra
 - The Rise of Digital Reading Virtual Chem Lab 5 Atomic Emission Spectra
 - Advantages of eBooks Over Traditional Books
2. Identifying Virtual Chem Lab 5 Atomic Emission Spectra
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Virtual Chem Lab 5 Atomic Emission Spectra
 - User-Friendly Interface
4. Exploring eBook Recommendations from Virtual Chem Lab 5 Atomic Emission Spectra
 - Personalized Recommendations
 - Virtual Chem Lab 5 Atomic Emission Spectra User Reviews and Ratings
 - Virtual Chem Lab 5 Atomic Emission Spectra and Bestseller Lists
5. Accessing Virtual Chem Lab 5 Atomic Emission Spectra Free and Paid eBooks
 - Virtual Chem Lab 5 Atomic Emission Spectra Public Domain eBooks
 - Virtual Chem Lab 5 Atomic Emission Spectra eBook Subscription Services
 - Virtual Chem Lab 5 Atomic Emission Spectra Budget-Friendly Options
6. Navigating Virtual Chem Lab 5 Atomic Emission Spectra eBook Formats

- ePub, PDF, MOBI, and More
 - Virtual Chem Lab 5 Atomic Emission Spectra Compatibility with Devices
 - Virtual Chem Lab 5 Atomic Emission Spectra Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Virtual Chem Lab 5 Atomic Emission Spectra
 - Highlighting and Note-Taking Virtual Chem Lab 5 Atomic Emission Spectra
 - Interactive Elements Virtual Chem Lab 5 Atomic Emission Spectra
 8. Staying Engaged with Virtual Chem Lab 5 Atomic Emission Spectra
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Virtual Chem Lab 5 Atomic Emission Spectra
 9. Balancing eBooks and Physical Books Virtual Chem Lab 5 Atomic Emission Spectra
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Virtual Chem Lab 5 Atomic Emission Spectra
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Virtual Chem Lab 5 Atomic Emission Spectra
 - Setting Reading Goals Virtual Chem Lab 5 Atomic Emission Spectra
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Virtual Chem Lab 5 Atomic Emission Spectra
 - Fact-Checking eBook Content of Virtual Chem Lab 5 Atomic Emission Spectra
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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