

Name: _____

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Activity Series Review Worksheet

Practice Using the Activity Series

Using the Activity Series Handout, circle the element that is more likely to replace the other in a compound.

Cu Sn or Ba Ni Al P or I Cr or Br Zn or

Predicting Single Displacement Reactions

Part 1: Write the chemical name for each compound below.

Part 2: Use the Activity Series to determine if a reaction will occur.

For those reactions that do occur, predict the products and balance the equation.
If no reaction occurs, write "NR" (No Reaction).

1. $\text{Mn} + \text{Cr}(\text{NO}_3)_3 \rightarrow$ $\text{Cr} + \text{Mn}(\text{NO}_3)_2$ assume
 manganese (II)
Manganese + Chromium (III) Nitrate → Chromium + Manganese (II) Nitrate
2. $\text{Cu} + \text{ZnSO}_4 \rightarrow$ NR assume copper
 (II)
Copper + Zinc Sulfate
3. $\text{H}_2 + \text{BiCl}_3 \rightarrow$ $2\text{Bi} + 2\text{HCl}$
Hydrogen + Bismuth Chloride → Bismuth + Hydrogen Chloride
4. $\text{F}_2 + 2\text{KCl} \rightarrow$ $\text{Cl}_2 + 2\text{KF}$
Fluorine + Potassium Chloride → Chlorine + Potassium Fluoride
5. $\text{Fe} + \text{CaSO}_4 \rightarrow$ NR assume iron
 (III)
Iron + Calcium sulfate
6. $\text{Pb} + \text{Zn}(\text{NO}_3)_2 \rightarrow$ NR assume
 lead (II)
Lead + Zinc nitrate
7. $\text{Co} + 2\text{H}_2\text{O} \rightarrow$ $\text{H}_2 + \text{Co}(\text{OH})_2$ assume cobalt (II)
Cobalt + Dihydrogen Monoxide → Hydrogen + Cobalt (II) Hydroxide
8. $\text{Zn} + \text{KNO}_3 \rightarrow$ NR
Zinc + Potassium nitrate
9. $\text{Ca} + \text{Co}(\text{NO}_3)_2 \rightarrow$ $\text{Co} + \text{Ca}(\text{NO}_3)_2$
Calcium + Cobalt(II) Nitrate → Cobalt + Calcium Nitrate
10. $\text{Sn} + \text{Pb}(\text{NO}_3)_2 \rightarrow$ $\text{Sn}(\text{NO}_3)_2 + \text{Pb}$ assume tin (II)
Tin + Lead (II) Nitrate → Tin (II) Nitrate + Lead
11. $\text{Mg} + 2\text{HNO}_3 \rightarrow$ $\text{H}_2 + \text{Mg}(\text{NO}_3)_2$
Magnesium + Hydrogen Nitrate → Hydrogen + Magnesium Nitrate

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1. Understanding the eBook The Activity Series Answer Key
 - The Rise of Digital Reading The Activity Series Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying The Activity Series Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an The Activity Series Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from The Activity Series Answer Key
 - Personalized Recommendations
 - The Activity Series Answer Key User Reviews and Ratings
 - The Activity Series Answer Key and Bestseller Lists
5. Accessing The Activity Series Answer Key Free and Paid eBooks
 - The Activity Series Answer Key Public Domain eBooks
 - The Activity Series Answer Key eBook Subscription Services
 - The Activity Series Answer Key Budget-Friendly Options

6. Navigating The Activity Series Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - The Activity Series Answer Key Compatibility with Devices
 - The Activity Series Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of The Activity Series Answer Key
 - Highlighting and Note-Taking The Activity Series Answer Key
 - Interactive Elements The Activity Series Answer Key
8. Staying Engaged with The Activity Series Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers The Activity Series Answer Key
9. Balancing eBooks and Physical Books The Activity Series Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection The Activity Series Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine The Activity Series Answer Key
 - Setting Reading Goals The Activity Series Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of The Activity Series Answer Key
 - Fact-Checking eBook Content of The Activity Series Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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