#### QUESTION 3

3. 1.Learners made a mini volcano in science laboratory by adding sodium bicarbonate to ethanoic acid. They added 100 ml of a 0, 2 mol.dm<sup>-3</sup> ethanoic acid solution to 10 g of NaHCO<sub>3</sub> to start the reaction of the volcano.

The balanced equation for this reaction is:

- 3.1.1 Define the term limiting reagent. (2)
- 3.1.2 Determine the limiting reagent in this reaction. (6)
- 3.1.3 Calculate the mass of the other substance in excess. (3)
- 3.1.4 Calculate the volume of CO<sub>2</sub> produced at STP. (4)
- 3.2 .A solution of potassium hydroxide (KOH) is prepared by dissolving 3,36 g crystals of KOH in 250 cm<sup>3</sup> of water.

Calculate the concentration of the potassium hydroxide solution.

(4)

3.3 .25 cm³ of a potassium hydroxide solution of concentration 0,25 mol.dm-3 completely neutralises a dilute solution of sulphuric acid (H<sub>2</sub>SO<sub>4</sub>) in a flask.

The unbalanced equation below represents the reaction that takes place:

$$H_2SO_4$$
 (aq) + KOH (aq)  $\rightarrow$  K<sub>2</sub> SO<sub>4</sub> (aq) + H<sub>2</sub>O ( $\ell$ )

- 3.3.1 Balance the above equation. (2)
- 3.3.2 Calculate the mass of sulphuric acid in the flask. (5)

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- 1. Understanding the eBook Taskmemo Chemistry 2015 Physical Science
  - The Rise of Digital Reading Taskmemo Chemistry 2015 Physical Science
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Taskmemo Chemistry 2015 Physical Science
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Taskmemo Chemistry 2015 Physical Science
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Taskmemo Chemistry 2015 Physical Science
  - Personalized Recommendations
  - Taskmemo Chemistry 2015 Physical Science User Reviews and Ratings
  - Taskmemo Chemistry 2015 Physical Science and Bestseller Lists
- 5. Accessing Taskmemo Chemistry 2015 Physical Science Free and Paid eBooks

- Taskmemo Chemistry 2015 Physical Science Public Domain eBooks
- Taskmemo Chemistry 2015 Physical Science eBook Subscription Services
- Taskmemo Chemistry 2015 Physical Science Budget-Friendly Options
- 6. Navigating Taskmemo Chemistry 2015 Physical Science eBook Formats
  - o ePub, PDF, MOBI, and More
  - Taskmemo Chemistry 2015 Physical Science Compatibility with Devices
  - Taskmemo Chemistry 2015 Physical Science Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Taskmemo Chemistry 2015 Physical Science
  - Highlighting and Note-Taking Taskmemo Chemistry 2015 Physical Science
  - Interactive Elements Taskmemo Chemistry 2015 Physical Science
- 8. Staying Engaged with Taskmemo Chemistry 2015 Physical Science
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Taskmemo Chemistry 2015 Physical Science
- 9. Balancing eBooks and Physical Books Taskmemo Chemistry 2015 Physical Science
  - Benefits of a Digital Library
  - o Creating a Diverse Reading Collection Taskmemo Chemistry 2015 Physical Science
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Taskmemo Chemistry 2015 Physical Science
  - Setting Reading Goals Taskmemo Chemistry 2015 Physical Science
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Taskmemo Chemistry 2015 Physical Science
  - Fact-Checking eBook Content of Taskmemo Chemistry 2015 Physical Science
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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