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$$\begin{array}{l} 1 \text{ N}_2 + \text{F}_2 \rightarrow 2 \text{ NF}_3 \\ 2 \text{ C}_6\text{H}_{10}\text{O}_2 + \text{CO}_2 + \text{H}_2\text{O} \rightarrow 3 \text{ HBr} + \text{KHCO}_3 + \text{H}_2\text{O} + \text{KBr} + \text{CO}_2 + \text{GaBr}_3 + \text{NaBr} + \text{SnO}_4 + \text{NaSO}_3 + \text{GaSO}_3 + 3 \text{ NF}_3 + 3 \text{ SnF}_2 + \text{N}_2\text{O}_3 \end{array}$$
 using the equation from problem 2 above answer the following

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$$\begin{array}{l} 1 \text{ N}_2 + 3 \text{ F}_2 \rightarrow 2 \text{ NF}_3 \\ 2 \text{ C}_6\text{H}_{10}\text{O}_2 + 17 \text{ O}_2 \rightarrow 12 \text{ CO}_2 + 10 \text{ H}_2\text{O} \\ 3 \text{ HBr} + 1 \text{ KHCO}_3 + 1 \text{ H}_2\text{O} + 1 \text{ KBr} + \text{CO}_2 + 2 \text{ GaBr}_3 + 3 \text{ Na}_2\text{SO}_3 + 1 \text{ Ga}_2\text{SO}_3 + 3 \text{ SO}_3 + 6 \text{ NaBr} + 3 \text{ SnO}_2 + 3 \text{ NF}_3 + 3 \text{ SnF}_2 + 1 \text{ N}_2\text{O}_3 \end{array}$$
 using the following equation $2 \text{ NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{ H}_2\text{O} + \text{Na}_2\text{SO}_4$

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$$\text{Ce} + \text{Zn} + \text{CuCl}_2 \rightarrow \text{ZnCl}_2 + \text{Cu} + \text{Zn}$$
 how many moles of Ce ZnCl_2 ZnCl_2 will be produced from 23.0 g of Ce Zn assuming Ce CuCl_2 CuCl_2 is available in excess

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$$\text{N}_2 + 3\text{F}_2 \rightarrow 2\text{NF}_3$$

$$2\text{C}_6\text{H}_{10} + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$$

$$3\text{HBr} + 1\text{KHCO}_3 \rightarrow \text{H}_2\text{O} + 1\text{KBr} + \text{CO}_2$$

$$2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 \rightarrow \text{Ga}_2\text{SO}_3 + 3\text{NaBr}$$

$$5\text{SNO}_2 + 3\text{NF}_3 \rightarrow \text{SNF}_2 + 2\text{N}_2\text{O}_3$$
 using the following equation

$$\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$$

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$$\text{C}_6\text{H}_{10} + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$$

$$3\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$$

$$2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 \rightarrow \text{Ga}_2\text{SO}_3 + 3\text{NaBr}$$

$$5\text{SNO}_2 + 3\text{NF}_3 \rightarrow \text{SNF}_2 + 2\text{N}_2\text{O}_3$$

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$$\text{SiO}_2 + 3\text{C} \rightarrow \text{SiC} + 2\text{CO}$$

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$$\text{N}_2 + 3\text{F}_2 \rightarrow 2\text{NF}_3$$

$$\text{C}_6\text{H}_{10} + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$$

$$3\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$$

$$2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 \rightarrow \text{Ga}_2\text{SO}_3 + 3\text{NaBr}$$
 solve the following stoichiometry grams grams problems 6

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web jun 30 2023 stoichiometry and balanced equations in stoichiometry balanced equations make it possible to compare different elements through the stoichiometric factor discussed earlier this is the mole ratio between two factors in a chemical reaction found through the ratio of stoichiometric coefficients

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web gabrielle m 9 years ago i m working on chemical reactions double and single replacement on flvs now my first question for this video is how do you have 4 aluminum atoms when it says 2Al subscript 2 do you just add the 2 and the subscript 2 same thing goes with the O_3 i am so lost

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web google classroom balance the following chemical equation $\text{Mg} + \text{OH}^- \rightarrow \text{Mg}(\text{OH})_2$ $\text{HCl} + \text{MgCl}_2$ H_2O note all reactants and products require a coefficient of at least one stuck review related articles videos or use a hint report a problem $7.4 \times 10^5 \theta \pi 8.52 \times 10^6$

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web a common type of stoichiometric relationship is the mole ratio which relates the amounts in moles of any two substances in a chemical reaction we can write a mole ratio for a pair of substances by looking at the coefficients in front of each species in

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