

SENIOR ONE (1) S.S. 1

CHEMISTRY

OBJECTIVE PARTS: SECTION A

TIME: 1 HOUR 45 MINUTES

ANSWER ALL QUESTIONS CORRECTLY

- 1.The aspect of chemistry that deals with the study of carbon and its compound is called (a) physical chemistry (b) Organic chemistry (c) Inorganic chemistry (d) Analytical chemistry
- 2.Which of the following is NOT an adverse effect of chemistry? (a) pollution (b) poisoning (c) emulsion (d) corrosion
- 3.A stage in scientific method of discoveries is (a) Evaporating (b) Hypothesis (c) Dehydrating (d) dissolving
- 4.A theory which has been extensively tested and proven correct without any exception becomes (a) scientific law (b) hypothesis (c) experiment (d) pattern
- 5.Career prospects in chemistry includes all the following EXCEPT (a) Biochemistry (b) Quality control (c) Accounting (d) Metallurgy
- 6.An apparatus in the laboratory in which the preparation of poisonous gases is done is called (a) distillation tank (b) Aspirator bottle (c) fume cupboard (d) bell jar
- 7.In the laboratory _____ is used for the collection of a gas over water (a) gas jar (b) Woulff's bottle (c) reagent bottle (d) burette
- 8.Two liquids which are immiscible with each other can be separated using (a) bell jar (b) gas jar (c) wash bottle (d) separating funnel
- 9.An apparatus used to hold drying agents during the preparation of dry gases is (a) conical flask (b) gas jar (c) Woulff's bottle (d) flat bottom flask
- 10.Coloured reagent bottles are used to store reagents which are capable of _____ easily by sunlight (a) evaporating (b) decomposing (c) dehydrating (d) dissolving
- 11.Which of the following changes is a chemical change? (a) melting of ice (b) liquefaction of air (c) slaking of lime (d) evaporation of a liquid
- 12.Which of the following substances will occupy a wider space? (a) carbon (iv) oxide (b) liquid milk (c) pieces of chalk (d) water
- 13.When a solid changes to gas directly, this process is called (a) freezing (b) sublimation (c) vaporization (d) evaporation
- 14.Which of the following changes produces a new substance ?(a) reaction of water with sodium chloride (b) addition of acid to base (c) turning of margarine to oil (d) evaporation of water
- 15.Which one of the following has a fixed shape and volume? (a) a cube of sugar (b) liquid wax (c) smoke (d) kerosene

Simple Questions And Answers For Chemistry Ss1

Kingsley Augustine



Simple Questions And Answers For Chemistry Ss1:

The Epistemology and Morality of Human Kinds Marion Godman, 2020-11-24 Natural kinds is a widely used and pivotal concept in philosophy the idea being that the classifications and taxonomies employed by science correspond to the real kinds in nature Natural kinds are often opposed to the idea of kinds in the human and social sciences which are typically seen as social constructions characterised by changing norms and resisting scientific reduction Yet human beings are also a subject of scientific study Does this mean humans fall into corresponding kinds of their own In *The Epistemology and Morality of Human Kinds* Marion Godman defends the idea of human kinds She first examines the scientific use and nature of human kinds considering the arguments of key philosophers whose work bears upon human kinds such as Ian Hacking John Searle Richard Boyd and Ruth Millikan Using the examples of gender ethnic minorities and Buddhism she then argues that human kinds are a result of ongoing historical reproduction chiefly due to pre existing cultural models and social learning Her novel argument shifts the focus away from the reductionism characteristic of research about human kinds Instead she argues that they are multiply projectable and deserving of scientific study not in spite of but because of their role in explaining our identity injustice and the emergence of group rights

College Chemistry MCQs Arshad Iqbal, 2017-08-29 College chemistry multiple choice questions has 1410 MCQs College chemistry quiz questions and answers MCQs on organic chemistry basic chemistry atomic structure chemical formulas chemical equations gas laws Charles's law Boyle's law inorganic chemistry MCQs with answers chemical science chemical reactions chemical bonding liquids and solids MCQs and quiz study guides for SAT ACT GAT GRE CLEP GED practice tests College chemistry multiple choice quiz questions and answers chemistry exam revision and study guide with practice tests for SAT ACT GAT GRE CLEP GED for online exam prep and interviews Chemistry interview questions and answers to ask to prepare and to study for jobs interviews and career MCQs with answer keys Experimental techniques quiz has 66 multiple choice questions Atomic structure quiz has 395 practice multiple choice questions Basic chemistry quiz has 73 multiple choice questions with answers Chemical bonding quiz has 166 multiple choice questions Gases and gas laws quiz has 241 multiple choice questions Liquids and solids quiz has 469 multiple choice questions Chemistry interview questions and answers MCQs on atomic mass atomic radii atomic radius absolute zero derivation Daltons law applications of Daltons law atomic absorption spectrum atomic emission spectrum periodic table electronegativity periodic table modern periodic table atomic spectrum atomic ionic and covalent radii atoms and molecules Avogadro number Avogadro's law azimuthal quantum number basic chemistry Bohr model Bohr's atomic model defects boiling point and external pressure boiling points bond formation Boyle's law charge to mass ratio of electron Charles's law chemical bonding chemical combinations chromatography classification of solids combustion analysis covalent radius covalent solids crystal lattice crystallization crystals and classification cubic close packing diamond structure diffusion and effusion dipole forces dipole induced dipole forces discovery of electron discovery of neutron discovery of proton dual

nature of matter dynamic equilibrium electron affinity electron charge electron distribution electron radius and energy derivation electron velocity electronic configuration of elements empirical formula energy changes and intermolecular attractions energy of revolving electron experimental techniques filter paper filtration crucibles fundamental particles gas laws gas properties graham s law grahams law of diffusion Heisenberg s uncertainty principle hexagonal close packing higher ionization energies hydrogen bonding hydrogen spectrum ideal gas constant ideal gas density ideality deviations intermolecular forces ionic radius ionization energies ionization energy isotopes kinetic interpretation of temperature kinetic molecular theory of gases Lewis concept liquefaction of gases liquid crystals liquids properties London dispersion forces magnetic quantum number mass of electron mass spectrometer metallic crystals properties metallic solids metals structure molar volume molecular ions molecular solids molecules moles Moseley law neutron properties non ideal behavior of gases orbital concept partial pressure calculations phase changes energies photons wave number Planck s quantum theory plasma state positive and negative ions pressure units properties of cathode rays covalent crystals properties of crystalline solids properties of positive rays quantum numbers quantum theory relative abundance Rutherford model of atom shapes of orbitals solid iodine structure solids properties solvent extraction spectrometer spin quantum number states of matter stoichiometry sublimation thermometry scales types of solids unit cell van der Waals equation vapor pressure and spectrum

Chemistry Jack Rudman,1980

General Chemistry Jack Rudman,National Learning Corporation,1979

Chemistry

Amin Elersawi,2014-07-15 This book helps students and readers visualize the three dimensional atomic and molecular structures that are the basis of chemical action An integral part of the text is to develop an explanation to hybridization which introduced to explain molecular structure when the valence bond theory failed to correctly envisage them Dr Elersawi presents the quantum theory of the electronic structure of atoms and focuses on the electronic structures and reactivity of atoms and molecules Many questions and answers of chemical components are introduced using molecular orbital and hybridization of orbitals The book has been made more informative and the subject matter has been presented in a very simple language clear style along with a large number of fully illustrative diagrams Atoms molecules ions chemical formulas and equations chemical bondings intermolecular forces energies electronegativity are offered to readers in effective and proven features clarity of writing and explanation If you are finding that Lewis dot structures are not enough for representing the atoms and molecules you are dealing with as a chemist then this is the book for you Overall this volume answers frequently asked questions and highlights the most important hybridized formulas It has a broader range than traditional quantum chemistry books It is a useful reference for health professionals practicing physicists chemists and materials scientists **Chemistry** Phillip John Buddle,J. S. Mackay,1970 Chemistry All-in-One For Dummies (+ Chapter Quizzes Online) Christopher R. Hren,John T. Moore,Peter J. Mikulecky,2022-11-23 Everything you need to crush chemistry with confidence Chemistry All in One For Dummies arms you with all the no nonsense how to content you ll need to pass your

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Student Guide For Living Chemistry
Barbara Ucko, 2012-12-02

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